

BASINKEVICH, I.R.

Literature on the woodpulp industry to be published in 1963.
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1. Starshiy redaktor Gosudarstvennogo izdatel'stva lesnoy i
bumazhnoy promyshlennosti.
(Bibliography—Woodpulp industry)

OLAKOWSKI, Tadeusz; KRZYSZKOWSKA, Anna; BASINSKA, Dobrosława; KAWCZYNSKA, Danuta; STANSKI, Wiktor.

Evaluation of BCG vaccination and its effectiveness in the Ostrołęka County. I. Evaluation of the status of BCG vaccination of children born in 1962. Przegl. epidem. 18 no.3:307-315 '64

Evaluation of BCG vaccination and its effectiveness in the Ostrołęka County. II. Tuberculin allergy in school children. Ibid.:317-324

1. z Wojewódzkiej Stacji Sanitarno-Epidemiologicznej w Aninie (dyrektor: dr. med. J. Zasławski); z Zakładu Epidemiologii Instytutu Gruzlicy (kierownik: doc. dr. O. Buraczewski); z Wojewódzkiej Przychodni Przeciwwirusowej (dyrektor: dr. med. J. Meissner) oraz z Potłatowej Przychodni Przeciwwirusowej w Ostrołęce (kierownik: lek. med. W. Stanski).

OLAKOWSKI, Tadeusz; KRZYSZKOWSKA, Anna; BASINSKA, Dobrosława; KAWCZYNSKA, Danuta; STANSKI, Wiktor.

Evaluation of BCG vaccination and its effectiveness in the Ostrołęka district. III. Postvaccination allergy in school children. Przegl. epidemiol. 19 no.1:31-37 '65

1. Z Wojewódzkiej Stacji Sanitarno-Epidemiologicznej w Aninie (Dyrektor: dr. med. J. Zasztowt); z Zakładu Epidemiologii Instytutu Gruźlicy (Kierownik: doc. dr. O. Buraczewski); z Wojewódzkiej Przychodni Przeciwgruźliczej (Dyrektor: dr. med. J. Meissner) i z Powiatowej Przychodni Przeciwgruźliczej w Ostrołęce.

1951

(6)

Problems of structural inorganic chemistry. J. K. Kozłowski
(Kopernik's Univ., Toruń, Poland). *Wiadomości
Chem.* 2, 178-82 (1948).--A review with 20 references.
Adam Sparyszak

CA

2

Use of heavy hydrogen and oxygen to explain the mechanism of chemical reactions. Helena Banińska (Copernik's Univ., Toruń, Poland). *Wzrost* 1974, 3, 113-18 (1979).--A review with 7 references. Adam Sporzynski

C. a.

/1951

(8)

The application of radioactive analysis in the studies of
meteorites. Helena Banitcka (Copernik's Univ., Toruń,
Poland). WISNOSTRZ CHYM. 3, 237-43 (1948).—A review
with 14 references. Adam Sporzyski.

CA

2

Ionic radii calculated from coagulation values for silver chloride hydroxide. *Hallin Baskin. Recueil Chem. 23, 260-7 (1940).* Radii of 10^{-8} cm. were calculated from crit. concns. of various cations coagulating primary particles. Ionic radii of K, Sr, Ba, La, Ce, and Th agree with those obtained by x-ray analysis of crystals. Ionic radii of H, Mg, Ca, Zn, Fe, and Al were found to be greater than those determined by x-ray analysis. This can be attributed to hydration of ions. The crit. concns. of coagulation for Pb^{++} and Zr^{4+} ions are lower than those for other ions of equal valency. *Sylvia Nowinski.*

BC

A-1
4

Stability of the hydrosol of silver chlorosulfate. II. Discussion
(Koenig, 1934, pp. 128-129). The stability of the hydrosol
of AgClS is determined by the change of light-absorption with time.
Over a period of 10 years, AgClS per se is not stable. The negative
sol has a light absorption of 0.1-0.2, while AgClS, the positive
sol, has a light absorption of 0.5-0.6. The positive sol is very
unstable and its light absorption decreases rapidly. The negative
sol, however, is very stable and its light absorption is never > 10%.
For the purpose of this study, AgClO₃ is added to KCNS in
the hydrosol. The hydrosol is much more transparent than when
AgClS is added to KCNS. This difference, however, disappears
after a while. The transparency of the negative sol is higher than that
of the positive sol in all cases, but with in-
creasing concentration of the electrolyte, but with in-
creasing concentration of the electrolyte, the regions in which the positive
sol is more transparent than the negative. For high concn. of
electrolyte the transparency increases owing to the formation of sol.
complexes between AgClS and KCNS. The change of trans-
parency with time shows that the coagulation is autocatalytic.
S. M. Rysicka.

BC

CONGULATION OF SILVER CHLORIDE HYDROCOL AS INTERPRETED BY OSTWALD'S THEORY. H. HASINAKA (KORRA. C.A.M., 1951, 26, 81-87).—The coagulation values of the primary AgCl particles are determined by using Lange's photocolormeter with two counterbalanced photocells. The activity coeff. of the coagulating ions are calculated from the Debye-Hückel formula, modified by Ostwald, viz. $\log f = \frac{\phi \cdot \sqrt{\mu}}{1 + \phi \cdot \sqrt{\mu}}$, where ϕ is a numerical coeff. depending on temp. and on the dielectric const. of the dispersing medium, μ = valency of the ion, μ = total ionic strength, and s = no. of ions of given species in the mol. The activity coeff. (f) thus calculated for ions of different valencies differ by amounts considerably exceeding the experimental error. Ostwald's rule, which requires constancy of f at the coagulation point, is therefore not strictly accurate. H. M. KRYNICKA.

A-1-7

ASD-11.4 METALLURGICAL LITERATURE CLASSIFICATION

SECTION	SUBSECTION	SECTION	SUBSECTION
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Basinska, H.

Chemical Abst.
Vol. 48 No. 8
Apr. 25, 1954
Inorganic Chemistry

New valences of iron, manganese, and chromium.
Halina Basinska (Kopernik Univ., Torun, Poland). *Wyd.
Zestawienie* 12-18 (1953).—The compds. of the follow-
ing metals are reviewed: Fe(IV), Mn(V), and Cr(IV), and
Cr(V). 11 references. Adam Sprawyński

Basinska, H.

✓ 33. New methods for the determination of cadmium. H. Basinska, *W. Chem. Zentr.*, 1958, 7 (10), 224. *References*, 22. *Chem.*, 1958, Abstr. No. 8730. — Volumetric and gravimetric methods for the determination of Cd are based on the interaction of Cd^{++} with $Li_2Fe(CN)_6$ to give a ppt of $Cd_2Fe(CN)_6$. (Tananaev and Kozlov, *Brit. Abstr. C*, 1951, 261). The methods are not applicable in the presence of metals that form slightly soluble ppt. with $Fe(CN)_6^{4-}$; salts of K and NH_4^+ with which Cd forms complex compounds of variable composition, also interfere. E. Hayes

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105T

BASINSKA, H.

Wiadomosci Chemiczne - Vol. 7, no. 9, Sept. 1953.

New valences of iron, manganese, and chrome. p. 412.

SO: Monthly list of East European Accessions, (EEAL), LC, Vol. 4, No. 9, Sept. 1955
Uncl.

Basinsak H.

Poland/ Analytical Chemistry - Analysis of Inorganic Substances

G-2

Abs Jour : Referat Zhur - Khimiya, No 3, 1957, 8446

Author : Basinsak, H. and Orylska, K.

Inst : Not given

Title : A New Method for the Titrimetric Determination of Cadmium

Orig Pub : Roczn. chem., 1956, Vol 30, No 1, 281-290 (in Polish with a summary in English)

Abstract : A volumetric method has been developed for the determination of Cd. The method consists in the direct titration of a standardized $\text{Li}_4\text{Fe}(\text{CN})_6$ solution with a solution prepared from the Cd salt to be investigated in the presence of indicators (sodium alizarin sulfonate, cresol purple, bromethymol blue). The method cannot be used when cations which form insoluble ferrocyanides (Zn, Cu, Pb) or potassium salts, forming complexes of varying composition with $\text{Cd}_2\text{Fe}(\text{CN})_6$ are present. For comparison purposes Cd was determined in $(\text{CH}_3\text{COO})_2\text{Cd}$, CdCl_2 , $\text{Cd}(\text{NO}_3)_2$ and CdBr_2 by the electrolytic method and by the method developed by the authors. When 0.1 M $\text{Li}_4\text{Fe}(\text{CN})_6$ portions not exceeding 20 ml are titrated with 0.25 M solutions of Cd salts at pH 5.35, the difference in the determinations by the two methods does not exceed 0.5%.

Card 1/1

-23-

Basinska, H.

POLAND / Analytical Chemistry. General.

E

Abs Jour: Ref Zhur-Khimiya, No 4, 1959, 11447.

Author : Basinska, H., Bobinska, T.

Inst : Not given.

Title : The Feasibility of the Application of Cupric Oxide as a Standard Reagent in Iodometry.

Orig Pub: Chem. analit., 1957, 2, No 3, 240-244.

Abstract: It was established that CuO, proposed by Rosenthal (RZhKhim, 1955, 581) to be used as a new standard reagent in iodometry, possesses a number of advantages in comparison with other standards: CuO is easily obtained in a pure form; is preserved in a glass vessel without change of its properties during many months; is not volatile or hygroscopic; besides, in the case of CuO, which

Card 1/3

POLAND / Analytical Chemistry. General.

E

Abs Jour: Ref Zhur-Khimiya, No 4, 1959, 11447.

Abstract: is obtainable usually in the form of a fine powder, exact weighing of required quantities is facilitated. To determine the titration standard of the $\text{Na}_2\text{S}_2\text{O}_3$ solution, about 0.2 g of CuO is weighed and placed in a conic flask; 10 ml of 2 n H_2SO_4 are added, and the mixture is heated over a hot water bath until CuO is dissolved completely; at cooling, the solution is diluted with water up to 50 ml, 1.5 g of solid KI are added, and the liberated iodine is titrated against $\text{Na}_2\text{S}_2\text{O}_3$ solution. Near the end of the titration period, 2 ml of a starch solution are added, and the titration is continued until the blue color disappears. In place of weighed portions of CuO , it is possible to apply solutions of CuSO_4 , pre-

Card 2/3

Country : POLAND
 Category : Analytical Chemistry. General Problems E
 Abs. Jour : Ref Zhur - Khim., No 5, 1959, No. 15040
 Author : Basinska, H.; Banasiak, K.
 Institut. :
 Title : Monohydrate of Barium Thiosulfate as a Standard Substance in Iodometry
 Orig. Pub. : Chem. analit., 1957, 2, No 4, 315-321
 Abstract : It was established that $\text{BaS}_2\text{O}_3 \cdot \text{H}_2\text{O}$ (I), suggested as a standard substance in iodometry (Ref Zhur-Khim, 1954, 15001), possesses a number of advantages in comparison with $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$ (II). I practically does not change its properties during storage in a glass jar for a long time. In addition, solutions of I differ considerably by their great stability in comparison with solutions of II. A shortcoming of I is its low solubility in water (2.66 g./l.
 Card: 1/3

Country	: POLAND	
Category	: Analytical Chemistry. General Problems	E
Abstr. Jour.	: Ref Zhur - Khim., No 5, 1959.	No. 15040
Author	:	
Institution	:	
Title	:	
Orig. Pub.	:	
Abstract	: at 25°), which provides no possibility for	
Cont'd	using solutions more concentrated than 0.01 M. However, the concentration of iodine can be easily and accurately determined by titration of separate weighed batches of I. In this case, about 100 ml. of water are added to a weighed batch of I, mixed well and titrated with an iodine solution. Towards the end of titration, 1 ml. of a starch solution is added. In the process of titration, I is gradually dissolved	

Card: 2/3

E - 9

BASINSKA, Halina; ORYLSKI, Zenon

Potentiometric titration of bismuth with potassium ferrocyanide.
Chem anal 4 no.4:685-689 '59. (HEAI 9:6)

1. Zaklad Chemii Nieorganicznej Uniwersytetu M.Kopernika, Torun.
(Potassium ferrocyanide) (Bismuth)

BASINSKA, Halina; PRZYBYSZEWSKA, Krystyna

New volumetric method of determining the potassium ferrocyanide with
o-dianisidine as an adsorption indicator. Chem anal 4 no.5/6:783-790
'59. (EBAI 9:9)

1. Katedra Chemii Nieorganicznej Uniwersytetu M.Kopernika, Torun.
(Potassium ferrocyanide)
(Dimethoxybenzidine)

BASINSKA, Halina; PRZYBYSZEWSKA, Krystyna

Volumetric determination of thorium by means of potassium ferrocyanide
in the presence of o- dianisidine. Chem anal 4 no.5/6:791-793 '59.
(EEAI 9:9)

(Thorium) (Potassium ferrocyanide)
(Dimethoxybenzidine)

BASINSKA, Halina; ORYLSKI, Zenon

Potentiometric titration of bismuth with sodium and lithium
ferrocyanide. Chem anal 5 no.2:187-191 '60. (REAT 10:3)

1. Katedra Chemii Nieorganicznej Uniwersytetu M.Kopernika, Torun.
(Bismuth) (Sodium ferrocyanides)
(Lithium ferrocyanide)

BASINSKA, Halina; BUKOWSKA, Alina

Volumetric determination of mercurous ions with alkaline ferrocyanide solutions in the presence of bromophenol blue as an adsorption indicator. Chem anal 5 no.3:355-360 '60.
(KEAI 10:8)

1. Katedra Chemii Nieorganicznej Uniwersytetu M. Kopernika, Torun.
(Mercury) (Ions) (Solutions) (Ferrocyanides)
(Bromophenol blue)

BASINSKA, Halina; BUKOWSKA, Alina

Volumetric determination of mercurous ions by means of potassium ferrocyanide in the presence of syndiphenylcarbazone. Chem anal 5 no.5:707-710 '60. (KEAI 10:9)

1. Katedra Chemii Nieorganicznej Uniwersytetu M. Kopernika, Torun.

(Mercury) (Ions) (Potassium ferrocyanide)
(Phenylazoformic acid phenylhydrazide)

BASINSKA, Halina; ORYLSKA, Krystyna

Volumetric determination of bismuth and thorium by means of potassium ferrocyanide in the presence of 3:3'-dimethyl-naphthidine. Chem anal 5 no.5:711-714 '60. (ZEAI 10:9)

1. Katedra Chemii Nieorganicznej Uniwersytetu M. Kopernika, Torun.

(Bismuth)	(Thorium)	(Potassium ferrocyanide)
	(Dimethylnaphthidine)	

BASINSKA, Halina; PRZYBYSZEWSKA, Krystyna

Volumetric determination of lead with alkaline ferrocyanide in the presence of 3:3'-dimethylnaphthidine. Chem anal 5 no.4:535-540 '60.
(EAI 10:9)

1. Department of Inorganic Chemistry, Copernicus University, Torun.

(Lead) (Ferrocyanides) (Dimethylnaphthidine)

BASINSKA, Halina; ORYLSKI, Zenon

On the solubility of bismuth ferricyanide $\text{Bi}[\text{Fe}(\text{CN})_6]$. Chem anal 6
no.3:307-308 '61.

1. Department of Inorganic Chemistry, Copernicus University, Torun.

BUKOWSKA, Alina; BASINSKA, Halina

Examinations of the composition of precipitates occurring in the
 $\text{Me}(\text{NO}_3)_2 + \text{Me}_4[\text{Fe}(\text{CN})_6]$ reaction where $\text{Me} = \text{K}, \text{Na}, \text{Li}$ in aqueous solu-
tion. Roczniki chemii 35 no.5:1211-1222 '61.

1. Department of General Chemistry, School of Agriculture, Olsztyn and
Department of Inorganic Chemistry, Copernicus University, Torun.

BASINSKA, Halina; ORYLSKI, Zenon; PERKOWSKI, Tadeusz

On the solubility of potassium-bismuth ferricyanide $\text{KBi}[\text{Fe}(\text{CN})_6]$.
Chem anal 7 no.5:911-914 '62.

1. Department of Inorganic Chemistry, N.Copernicus University, Torun.

BASINSKA, Halina; ORYLSKA, Krystyna; MURAWSKA, Zofia

Indirect manganometric method of determining bismuth by precipitation with oxalic acid. Chem anal 8 no.2:151-155 '63.

1. Department of Inorganic Chemistry, M.Copernicus University, Torun.

BASINSKA, Halina; ORYLSKI, Zenon; CHARASZKIEWICZ, Aleksandra

Solubility of bismuth ferrocyanide $\text{Bi}_4 [\text{Fe} (\text{CN})_6]_3$.
Chem anal 8 no.3:473-474 '63.

1. Katedra Chemii Nieorganicznej, Uniwersytet M. Kopernika,
Toruń.

BASINSKA, Halina; MILKE, Helena; ORYLSKI, Zenon; RYCHLIK, Wieslaw

Studies on the adaptation of alkali ferrocyanide and ferricyanide in volumetric analysis. *Studia Tor chemia* 5 no. 1: 67-77 '64.

1. Department of Inorganic Chemistry, N. Copernicus University, Torun, and Department of General Chemistry, School of Agriculture, Olsztyn.

SARNOWSKI, Marcin; BASINSKA, Helena; LUKASZYK, Andrzej

Influence of local application of chloramphenicol in the course of the healing of noninfected wounds in rabbits. Pol. przegl. chir. 37 no.2:159-167 F '65.

1. Z III Kliniki Chirurgicznej Akademii Medycznej w Poznaniu (Kierownik: prof. dr. A. Piskorz) i z Zakładu Histologii Prawidłowej i Embriologii Akademii Medycznej w Poznaniu (Kierownik: prof. dr. K. Mietkiewski).

BASINSKA, Maria; RUTKOWSKI, Wladyslaw

Comparison of distillation and extraction methods of separating trace amounts of germanium; application for the determination of germanium in minerals and coal. Chem anal 8 no.3:353-360 '63.

1. Department of Analytical Chemistry, Institute of Nuclear Research, Warsaw.

~~PAMPUCHOWA~~
BASINSKA-PAMPUCHOWA, ST.

✓ Application of thermal differential analysis for the investigation of clays of the Strzegom deposits. St. Basinska-Pampuchowa and W. Szymborski. *Prace Inst. Mineralogii Halic. 7, 276-83 (1955) (English summary)*.—All the clays investigated are used as refractories. Any change in the thermal differential curve found was rechecked by aid of x-rayograms in order to find out what the breaks in the curve signified. The clays investigated were a kaolin, a montmorillonite, a bentonite, a halloysite, and a kaolin with 5% peat, as this mixt. occurs naturally. In this way it was found that the presence of montmorillonite in any clay will cause an endothermic change at 150°, which otherwise is absent. Another endothermic change is noted at 600-620°. This corresponds to crystallographic changes of kaolinite; an endothermic reaction is noted at 870-920°, which signifies either the presence of γ - Al_2O_3 or the formation of corundum.
Werner Jacobson

(U. S. GOVERNMENT PRINTING OFFICE)		(U. S. GOVERNMENT PRINTING OFFICE)	
PROCESSES AND PROPERTIES		PROCESSES AND PROPERTIES	
BC 2-1 The rate of dissolution of metals in aqueous acid solutions. H. R. HARRIS and A. BARSZAKI (Trans. Chem. Soc. 1934, 56, 71-83).—The process of dissolution of metals in acidic chloride solution is a complex one. In general terms, diffusion plays only a minor role in this system, as is shown by the experimental results in the velocity of reaction of metals, and by the temperature coefficients of the reaction (10-12). The addition of hydrochloric acid or of chloride anion to the solution, as a result of the formation of complexes of metallic chloride. The velocity coefficient of the reaction increases rapidly with increase in concentration of anions and hydrogen chloride, whence it is concluded that the concentration of the HCl formed from hydrogen or anionous chlorides, determines the velocity of reaction. R. TAPKOWSKI.			
A. S. S. A. METALLURGICAL LITERATURE CLASSIFICATION		METALLURGY	
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SOME PROPERTIES		SOME PROPERTIES	

1ST AND 2ND OBJECTS												3RD AND 4TH OBJECTS											
PROCESSES AND PROPERTIES INDEX																							
<i>VN</i> Velocity of solution of metals in bromine water. A. Hasiński. <i>Roczniki chem.</i> 12, 530-50(580) in German)(1932).—The rates of soln. of Cu, Zn and Ni in aq. Br in presence of KBr are rather high and identical, when detd. at 10°, 15°, 20° and 25° at similar conditions of stirring. The soln. is a diffusion process, having a temp. coeff. for the interval 15-25° in all cases 1.43. Soln. of Sn and Cu proceeds in the following phases: $\begin{array}{l} \text{Sn} + \text{Br}_2 \rightarrow \text{SnBr}_2 \xrightarrow{\text{Br}_2} \text{SnBr}_4 \\ \text{Cu} + \text{Br}_2 \rightarrow \text{CuBr}_2 \xrightarrow{\text{Br}_2} \text{CuBr}_4 \end{array}$ and 2Cu + Br₂ → 2CuBr; and 2Cu + Br₂ → 2CuBr → 2CuI, → 2CuI₂, and the second in soln. The rate of the process is expressed by the formula: $vdx/dt = k_1(a-x) + k_2x$, where k_1 and k_2 are the velocity coeffs. of the first and the third reactions, resp., x is the vol. of aq. Br in cc., S the area of the Sn and Cu surfaces in sq. cm., a the initial Br concn. and x the concn. of the reaction products after t min. J. W.																							
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variation of electrokinetic potential of silver halide colloids during osmosis.
 ANTONI. RANOWSKI. *Rozprawy Chem.* 13, 117-24 (in German 124) (1931), cf. C. A. 27, 217. Ultramicroscopic investigations in neg. and pos. Ag halide colloidal sols (AgCl, AgBr, AgI) led to the following results: (1) The electrokinetic potential of a positively charged sol decreases with increased diln., passes zero and in great dilns. becomes neg. (2) When colloidal AgI is dild. with ultrafiltrate the potential of the particles decreases, because of coagulation, without changing sign. FELIX P. LUK

430 554 DETAIL LITERATURE CLASSIFICATION

777 PROCESSING AND PROPERTIES INDEX TOP AND 1TH COPIES									
*The Rate of Dissolution of Copper in Copper Salt Solutions. Antoni Hualski. (Rozmiki Chem., 1934, 14, 31-44; Chem. Zvest., 1934, 186, 1, 3331).—The effects of temperature (10°-35° C.), rate of stirring (300-500 r.p.m.), and concentration of the solution on the rate of dissolution of copper in copper ammonium chloride and bromide solutions have been determined. The dissolution is a diffusion phenomenon of the first order, the reaction constant increasing by about 30% for a rise in temperature of 15° C. (10°-25° C.). Passage of carbon dioxide through the solution reduces the rate of reaction slightly, but changes in the concentration of copper or of ammonium chloride have little effect on the reaction constant.—A. R. I.									
AS 111.1 METALLURGICAL LITERATURE CLASSIFICATION									
1000 111.1.1 10000 111.1.1.1 10000 111.1.1.1.1 10000 111.1.1.1.1.1									

JST AND JND CODES		PROCESSING AND PROPERTY INDEX		JST AND JND CODES	
ca		Stability of colloids. Stability of negatively charged colloidal silver iodide. A. Basinski. <i>Recueil Chem.</i> 36, 1017-37(1934).—The stabilizing action (I) of halides is, up to a certain limiting concn. (II), independent of the nature of the cation; at higher concns. I diminishes in the order Li^+, Na^+, K^+, NH_4^+. The value of II for different anions diminishes in the order Cl^-, Br^-, I^-. A theoretical interpretation is advanced. B. C. A.		2	
ADD-514 METALLURGICAL LITERATURE		CLASSIFICATION		8-2-1-1-1-1-1-1	
FROM SYNDICATE		FROM SYNDICATE		FROM SYNDICATE	

[illegible]

11/ APR 1968		PROCESSED AND REPROCESSED		APR 1968	
BC Stability of colloids. II. Negative silver iodide colloid. A. Hammett (Ham. Chem., 1933, 15, 430-440; cf. A., 1933, 237).—The coagulating action of cations on dialyzed negatively charged AgI sol diminishes in the series $Li > Na > K > NH_4 > H > Ag$; $Ca > Mg > Mn^{2+} > Co > Ba > Ni > Cu > Sr > Cu > Ba > Pb^{2+}$; $Fe^{3+} > La^{3+} > Ce^{3+} > Al$; $Zr^{4+} > Th^{4+}$. The action of univalent cations is antagonistic to that of bi- and tri-valent ions; when uni- and quadri-valent ions are present together antagonism is observed only with $LiCl$. An interpretation of the above phenomena, based chiefly on considerations of exchange adsorption, is advanced. R. T.					
ABSTRACT METALLURGICAL LITERATURE CLASSIFICATION					
SOURCE SYNDICATE		SOURCE HIG GUY SOC		SOURCE BOWLING	
SOURCE OF		SOURCE		SOURCE OF GUY 111	

COMMON ELEMENTS		PROCESSES AND PROPERTIES INDEX	
<i>2</i> The theory of adsorptive solution. A. Rasinski. <i>Rozprawy Chem.</i> 17, 567-77 (in English 576-77) (1937). The adsorption theory of soln. of a solid in a liquid is dis- cussed and the phenomena accompanying the soln. are pointed out and characterized. M. Wojciechowski.		<i>2</i>	
ASB-SLA METALLURGICAL LITERATURE CLASSIFICATION		1930-1949	
1950-1959		1960-1969	
1970-1979		1980-1989	
1990-1999		2000-2009	

1ST AND 2ND SECTIONS		3RD AND 4TH SECTIONS	
PROCESSES AND PROPERTIES INDEX			
<i>A</i>		<i>A-1</i>	
Coagulation of silver iodide hydrosol. A. BARNETT (Recs. Chem., 1938, 18, 335-349).—Measurements of the activities of ions at various concns. causing flocculation of AgI hydrosols of various concns. at different temp., and in presence of different concns. of EtOH, do not confirm Ostwald's hypothesis (A., 1937, 614 and preceding abstracts), except as a special case. R. T.			
ASD-SLA METALLURGICAL LITERATURE CLASSIFICATION			
FROM SYNONYM		FROM BROWSE	
100000 100 1000 10000 100000		10000 100 1000 10000 100000	

2 Stability of colloidal silver halides: Pure silver bromide sol. Ansoni Borkh. Recomb. Chem. 21, 59-65 (1947).—The product is AgBr sol of high concn. and purity is described. The sols were prepared by electrolysis and coated by electrodeposition, and contd. (per l.) 280 millimol. AgBr, $1-8 \times 10^{-3}$ g. ions Br, and $3-4 \times 10^{-3}$ g. ions H. The sol had the character of a strong electrolyte, since 90-100% of the protons were electrometrically active. The AgBr sol was relatively stable in spite of the small elec. surface charge. H. H. Semant	
ASS-516 METALLURGICAL LITERATURE CLASSIFICATION	EDMONT 501137
EDMONT 501137	EDMONT 501137

C. a.

1951

"The application of radioactive indicators in chemical analysis. Antoni Boniński (Kopernik's Univ., Toruń, Poland). *Wiskomosci Chem.* 7, 222-30 (1949).—A review with 22 references. Adam Sporyński

CA

18

Rate of vaporization of brines. Anton Basthelt and Zenon Casiminski. *Proceedings of the 8th International Conference on the Recovery of NaCl from Salt, Soins, of rock salt* (NaCl 97.81, KCl 0.87, MgCl₂ 6.18, MgSO₄ 0.87, CaSO₄ 0.06, and Ca(HCO₃)₂ 0.08%), evapn. was greatly retarded owing to formation of a supernatant crust of pyramid like crystals of NaCl. This was caused by river water used for making up the brine and having a temporary hardness of 8.16-19.30° German. The greater the hardness, the tougher the crust. With distd. H₂O, evapn. was rapid, the surface was free from crust, and the crystals were cubical throughout. The following means were effective in overcoming the deleterious effect of Ca(HCO₃)₂ in the brine: (1) addn. of org. or inorg. acids in amt. sufficient to reduce the pH to 3.10-3.50; (2) addn. of various vegetable oils (which, however, form insol. Ca soaps), paraffin oil, paraffin, or turpentine to provide 0.238-0.370 mg. of the surface-tension-reducing agent per each sq. cm. of brine surface; (3) filtering through adsorbents such as activated C, silica gel, "kationite," or Permutit; (4) addn. of multivalent metal salts ThCl₄, ZrO₂·3H₂O, CeCl₄, AlCl₃, La(CH₃COO)₃, UO₂(NO₃)₂, KAl(SO₄)₂·12H₂O, FeCl₃, and "didymium" nitrate, which undergo hydrolysis to form free acids for the reaction with Ca(HCO₃)₂, whereby CO₂ is liberated and the metals are converted to gelatinous hydroxides acting probably as adsorbents; (5) addn. of K₂Fe(CN)₆ or K₃Fe(CN)₆, which, however, causes crystn. of NaCl in the form of fine powder having a pronounced tendency to creep upwardly along the wall of the glass vessel; (6) addn. of (NH₄)₂C₂O₄ or (NH₄)₂CO₃ followed by removal of the ppt. in the latter case; and (7) electrolysis with current of 2.4 v. and c.d. of 0.5 milliamp./sq. cm., which gave very favorable results on an industrial scale. Expts. with artificial brines showed that Ca and Mg ions in combination with SO₄²⁻ cause crystn. of NaCl in the form of hollow pyramids which are buoyed by air and steam but readily sink, unless Ca(HCO₃)₂ is present in relatively large amt. to cause their agglomeration into a compact crust. Since 1948, KAl(SO₄)₂ has been used on an industrial scale with a large increase in output. Bruno C. Metzner

CA

Radioactive carbon isotopes. Antoni Bialicki (Kopernik's Univ., Torun, Poland). *Wydawnictwo Chem. 4, 75-80 (1980).* - A review with 31 references. A. Szwarc

Pasinski, H.

on the rate of action of a photographic developer containing metol and quinol
is studied by following the changes in its redox potential with time. Of the ten
substances examined ascorbic acid, caffeine and stannous chloride are the most

WGS

BASINSKI, A.

Polish Technical Abst.
No. 1 1954

Chemistry and Chemical Technology

Basinski A., Jakubowski S. Nowe Metody of Quantitative Determination of Cadmium. Part I.

„Nowe metody ilościowego oznaczania kadmu”. Przemysł Chemiczny No. 4, 1953, pp. 190—193, 4 tabs.

A quick and exact gravimetric method of determining cadmium by precipitation with a small excess of lithium ferrocyanide from a neutral solution in the form of $Cd_2Fe(CN)_6$. The difference in determinations does not, as compared with the electrolytic method, exceed 0.2%. The method cannot be applied in the presence of cations (zinc, copper, lead) which, with ferrocyanide anion, form compounds not readily soluble; it can neither be used in the presence of potassium and ammonium salts, which with cadmium ferrocyanide produce complex compounds of variable composition.

NASINSKI, A.

Part II. Titrimetric determination of cadmium by precipitation in the form of ferrocyanide.

"Nowe metody ilościowego oznaczania kadmu. II. Miareczkowe oznaczanie kadmu przez wytrącenie w postaci żelazocyanku". *Przemysł Chemiczny*, No. 5, 1953, pp. 341-342, 4 tabs.

A method is described of titrimetric determination of cadmium ions by precipitation from neutral solution such as $Cd_2Fe(CN)_6$ with a determined excess of lithium ferrocyanide standard solution, and by titration of this excess with potassium permanganate. It was established that the method is accurate and suitable for practical use. The method

cannot be used in the presence of cations which form slightly soluble ferrocyanides (such as zinc, lead, copper), or in the presence of potassium and ammonium salts. The determination should be carried out in neutral solutions and in the absence of oxidizing compounds.

Basinski, A.

548 : 541.2

2023

• Basile's A. Inorganic and General Chemistry, Vol. 2

* Basile's A. Inorganic and General Chemistry, Vol. 2
„Chemia nieorganiczna i ogólna” Tom 2. Warszawa, 1953. PWN. 16.

303 22 112 115 62 100

192 pp. 112 figs. 62 tabs.
Boron. Colloids. Periodic system of chemical elements. Transuranic elements. Metals. Electrochemistry. Solubility product and activity product. Alkali metals. Copper group. Alkaline earth metals. Radioactive elements. Elementary molecules and atomic nuclei. Isotopes. Nuclear transformations by artificial means. Importance of artificial radioactivity. Utilization of nuclear energy. Photoelectric effect. Compton's effect. Activation potential. Ionization energy. Zinc group. Aluminum group. Rare earth elements. Heavy metals of silver and germanium subgroup. Titanium subgroup. Vanadium subgroup. Manganese subgroup. Manganese subgroup. Iron group. Platinum metals. Complex compounds.

BAEINSKI, A

POL.

8804

841.183.25 : 545.224-35 : 001.183.3

Bosiński A., Kaczmarek T. Sorption of Sulphurous Acid by Potato Starch.

„Sorpceja kwasu siarkowego przez mączkę ziemniaczaną”. Przemysł Chemiczny. No. 1, 1934, pp. 23--28, 12 rys., 2 tab.

The influence was investigated of concentration, temperature and acidity of solution on the sorption of sulphurous acid by potato starch in water solution. The sorption increases with rise in temperature, the rate of increase depending on the concentration of sulphurous acid in solution. The relationship between the concentration of sulphurous acid in solution and the extent of sorption can be presented approximately by means of a linear function. When the pH of the solution increases, the sorption decreases, attains minimum at the pH range of from 4 to 6, then increases till the pH value reaches 10; at any higher pH the extent of sorption decreases slowly.

①

BASINSKI, A.

POL

081.182.3 : 541.163.24

Basinski A., Wilmaniska A. Influence of Electrolytes on the Sorption of Sulphurous Acid by Potato Starch.

"Wplyw elektrolitów na sorpcję kwasu siarkowego przez mączkę ziemniaczaną". Prace Instytutu Chemii, No. 2, 1934, pp. 82-84, 3 figs., 5 tabs.

The influence was investigated of electrolytes, NH_4Cl , NaCl , KCl , LiCl , SrCl_2 , CaCl_2 , and MgCl_2 , on the sorption of sulphurous acid by potato starch in an aqueous solution. It was established that the presence of these electrolytes raises the sorption, the increase being contingent on the increased concentration. No distinct differences of action of individual ions in a 0.25 N electrolyte solution were observed. In 0.5 N solution, the progressively decreasing influence of ions can be represented as follows: $\text{Ca}^{++} > \text{Mg}^{++} > \text{Sr}^{++} > \text{Li}^{+} > \text{K}^{+} > \text{Na}^{+} > \text{NH}_4^{+}$. In higher electrolyte concentrations the influence of magnesium ions is stronger, i.e., that of calcium ions. Glycerin and alcohol raise the sorption rate, and the influence of these substances increases with the degree of concentration.

①

BASINSKI, A.; KACZARSKA, T.

"Adsorption of Sulfurous Acid by Potato Starch." P. 23, (PRZEMYSŁ CHEMICZNY, Vol. 10, No. 1, Jan. 1954, Warszawa, Poland)

SO; Monthly List of East European Accessions, (EEAL), LC, Vol. 4, No. 1, Jan. 1955 Uncl.

BASINSKI, A.

Basinski, A. [unclear] M. New Methods of Quantitative Determination of Cadmium. III. Potentiometric Titration of Cadmium with Standard Lithium Ferrocyanide Solution.

Nowe metody ilościowego oznaczania kadmu. III. Potencjometryczne miareczkowanie kadmu żelazocyjankiem litowym". Przemysł Chemiczny. No. 3, 1955, pp. 145-147, 3 figs., 2 tabs.

A description of a potentiometric method of determining cadmium ions by titration at room temperature with standard lithium ferrocyanide solution. The method is rapid and accurate, the error in determination not exceeding 0.2%. It is not possible to use the above method in the presence of those cations which, with ferrocyanide, form slightly soluble salts (as zinc, lead, copper) and in the presence of potassium ions, with which cadmium ferrocyanide forms complex compounds of varying composition.

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Basinski, H.

4

corresponding value. Negative results were obtained in the range of concentrations from 10⁻⁴ to 10⁻² M CdCl₂. When cadmium acetate was used, the results were positive but only in a determined range of concentrations.

(Clipped Abstract)

CH/ep
August 10, 1956

RM
JH

Distr: 4E20 27

✓ Determination of cadmium. III. Potentiometric titration of cadmium by aid of lithium ferrocyanide. A. Basileki and M. Lango (M. Kopernikus Univ., Torun, Poland). *Prace Chem.* 11(34), 145-7 (1955); *Ch. C.A.* 49, 778a. The titration of Cd^{2+} at room temp. with a 0.2423M $Li_2Fe(CN)_6$ soln. can be done smoothly at pH 5-7, and any cations will disturb which will form slightly sol. salts with the $Fe(CN)_6^{4-}$ like Zn, Pb, or Cu. K ions too must be absent. The error is $\pm 0.5\%$. Werner Jacobson.

12-25-1961, A.

Distr: 4E20(j)/4E2e

✓ Determination of cadmium. IV. Conductometric titration of cadmium with lithium ferrocyanide. A. Basinski and Z. Orylski. *Przemysl Chem.* 11(34), 203-4 (1966), cf. preceding abstr. The following salts were analyzed: CdSO_4 , $\text{Cd}(\text{NO}_3)_2$, CdCl_2 , and $\text{Cd}(\text{OAc})_2$. The current used was of 2000 cycles/sec. The semimicroburet was of 5-ml. vol., with a scale of 0.01 ml. The design of the app. is given. The voltage used was 0.5-0.6, resistance 80-130 ohms. Thus: 5 or 10 ml. of $\text{Li}_2\text{Fe}(\text{CN})_6$ (I) (0.2974N) were dild. to about 50 ml., and titrated with the Cd salt of a concn. several times higher than I. For CdSO_4 and $\text{Cd}(\text{OAc})_2$ the cond. had a marked min. at the equil. point; for $\text{Cd}(\text{NO}_3)_2$ and CdCl_2 the cond. was increasing all the time, but the rate changed markedly at the equil. point. The results were compared with an electrolytic method; the differences were $\leq 1\%$. The method is not applicable in presence of cations which form slightly sol. ferrocyanides, and with salts giving complex compds. with $\text{CdFe}(\text{CN})_6$.
L. G. Mantina

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BASINSKI, A.

chem
 ✓ Not formation of ferric, aluminum and chromium hydroxides with ion-exchange. A. Basinski and M. Sierocka (Roczn. Chem. 1956, 30, 666-683). The formation kinetics of Fe, Al, and Cr hydroxide hydrosols, prepared by treatment of the metal chlorides (as solutions containing 0.1 mol. of metal chloride and 0.15 mol. of HCl in 1 l.) with a basic ion-exchanger (Wofatit MD treated with 4% NaOH), are investigated by adding 0.5-1 g. portions of the anionite to 50-ml. samples of the metal chloride solutions to the point of colloid-coagulation and measuring the pH and electrolytic conductivity in the clarified solutions at 10-min. intervals during the operation. The relation between conductivity, pH of the solution and quantity of added ion-exchanger during the course of neutralization is shown in graphs. The dependence of pH and conductivity values from the extent of neutralization of FeCl₃ and the dependence of conductivity on the pH of this solution show typical decline in conductivity and pH values during the process of neutralization is explained by the formation of Fe(OH)₃ hydrosol. The resulting hydrosol is stable at pH between 2.00 and 4.20. The process of neutralization of AlCl₃ is shifted 2-40 pH units towards the alkaline range and occurs at pH between 5.5 and 4-48. In the case of CrCl₃, neutralization begins above pH 3.50 and the functional independence between pH-values, conductivity and quantity of added anionite is rather indistinct.

L. S.

PM

~~SECRET~~ (ANTON)

7 21 7
~~Reaction of iron(II) with potassium ferri-
cyanide (K₃Fe(CN)₆) and Winkler's reagent (K₂Cr₂O₇)~~
~~by J. (Anton, Poland). J. Chem. Soc. 1918-19~~
~~1919 (1919-1920). The method consists in the~~
~~direct titration of Cd²⁺ with standard K₃Fe(CN)₆ soln.~~
~~in presence of (NH₄)₂SO₄. Under these conditions Cd₂K₂~~
~~[Fe(CN)₆] is formed in the ratio Cd²⁺: [Fe(CN)₆]³⁻ =~~
~~1.25. The titration must be carried out in neutral or~~
~~slightly acid soln. and in absence of cations yielding slightly~~
~~sol. salts with [Fe(CN)₆]³⁻ anion, such as Zn, Pb, or Cu.~~
P. M. B. 2

NT

Antoni Basinski and Witold Bogotko: "Potentiometric Titration of Cadmium with Potassium Ferrocyanide, "Roczniki Chemii, Vol 30, No 2, Warsaw, 1956. Published from the Research Laboratory of Physical Chemistry, M. Kopernik University, Torun, 15 Sep 55.

POLAND/Physical Chemistry - Solutions, Theory of Acids and Bases.

B-11

Abs Jour: Referat. Zhurnal Khimii, No 2, 1958, 3947.

Author : Antoni Basinski, Maria Koneczna.

Inst :

Title : Solubility of Silver Pyrophosphate in Alkali Pyrophosphates.

Orig Pub: Roczn. chem., 1956, 30, No 4, 1077-1082.

Abstract: The solubility of $\text{Ag}_4\text{P}_2\text{O}_7$ (I) in the solution of $\text{Na}_4\text{P}_2\text{O}_7$ (II) of various concentration was investigated at 20, 30, 40, 50 and 60°. It was revealed that the solubility rises with the II concentration. The solubility of I in II solutions rises with the temperature rise up to 50°, after which it drops. It was found that the solubility of I in $\text{K}_4\text{P}_2\text{O}_7$ solutions is considerably higher than in equimolar II solutions.

Card : 1/1

-15-

Basinski, Antoni

Zarys fizykochemii koloidow.

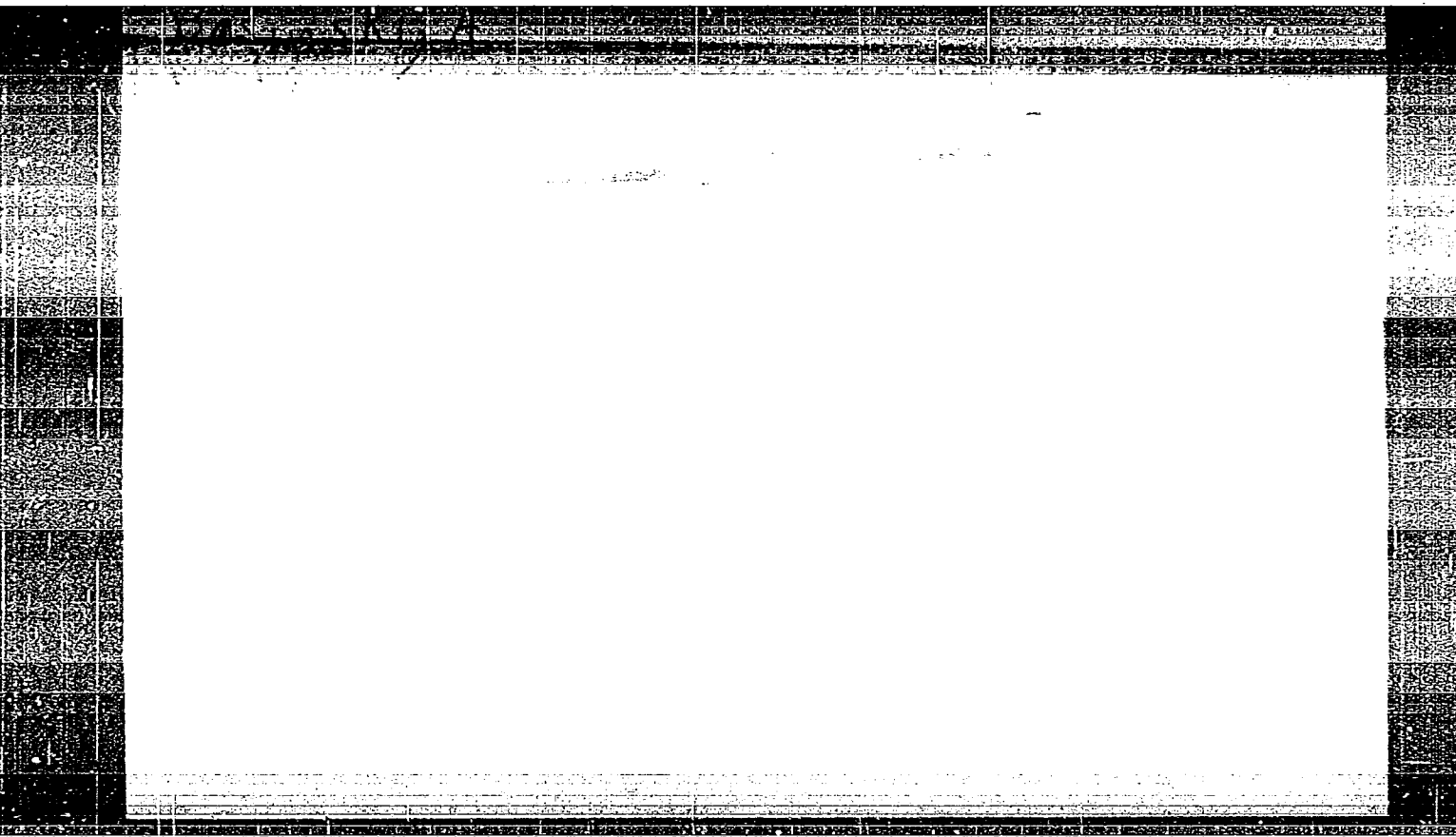
Warszawa, Poland; Panstwowe Wydawn, Naukowe, 1957., 492 P

Monthly List of East European Accessions (EEAI) LC, Vol. 9, no.2, Feb. 1960

Uncl

"APPROVED FOR RELEASE: 06/06/2000

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APPROVED FOR RELEASE: 06/06/2000

CIA-RDP86-00513R000203910002-4"

BASINSKI, ANTONI

POLAND/Physical Chemistry - Colloid Chemistry. Disperse Systems.

B-14

Abs Jour : Ref Zhur - Khimiya, No 8, 1958, 24375

Author : Basinski Antoni, Sierocka Michalina, Schmidt Felicja

Inst : -

Title : Mechanism of Preparation of Sols of Molybdic Acid by Means of Ion-Exchange Resins.

Orig Pub : Roczn. chem., 1957, 31, No 1, 3-11

Abstract : Sol of molybdic acid (I) was obtained by ion-exchange between ammonium molybdate and Wofatit F. The course of exchange adsorption was investigated by determinations of the pH value and conductivity of the solution. It was found that sol of I is formed at pH 2.0-1.6.

Card 1/1

14

APPROVED FOR RELEASE: 06/06/2000

CIA-RDP86-00513R000203910002-4"

POLAND/Chemical Technology - Chemical Products and Their Application - Photographic Materials.

Abs Jour : Ref Zhur - Khimiya, No 9, 1958, 30042

Author : Basinski, A., Tomczyk, M., Zielinski, L.

Inst : -

Title : Investigations on the Stabilization of Photographic Developers.

Orig Pub : Roczniki Chem, 31, No 1, 165-173 (1957) (in Polish with an English summary)

Abstract : The effect of various substances on the shelf-stability of anydol, pyrogallol, glycine, and metolhydroquinone developers has been investigated by sensitometric and by potentiometric methods. Lactic acid was found to double the useful life of the first three of the above developers. Anydol is stabilized by the addition of hydrocarbons, metol, caproic acid, and by acetone sulfite. Glycine can be stabilized with hydrocarbons,

Card 1/2

POLAND/Chemical Technology - Chemical Products and Their
Application - Photographic Materials.

H.

Abs Jour : Ref Zhur - Khimiya, No 9, 1958, 30042

pyrogallol, hydroquinone, and by the addition of small amounts (0.04 gms/liter) of semicarbazine hydrochloride and hydrazine sulfate. The shelf-stability of metolhydroquinone developers can be extended by more than 20% by the addition of succinic, oleic, and lactic acids as well as by the addition of acetone sulfite and tertiary butanol. Tannin and cysteine reduce the shelf-stability of the developers tested by 30%.

Card 2/2

46

BASINSKI, A.

SCIENCE

Periodicals: ROCZNIKI CHEMII. Vol. 31, no. 2, 1957.

BASINSKI, A. The solubility of cadmium ferrocyanide in water. p. 457.

Monthly List of East European Accessions (ERAI) LC, Vol. 8, No. 4,
April 1959, Unclass.

BASINSKI, A.

POLAND / Inorganic Chemistry, Complex Compounds.

C

Abs Jour: Ref Zhur-Khimiya, No 18, 1958, 60424.

Author : ~~Antoni Basinski~~, Mieczyslaw Skrzypek.

Inst : -

Title : Reaction Course in $\text{MnSO}_4 + \text{Li}_4[\text{Fe}(\text{CN})_6]$ System.

Orig Pub: Roczn. chem., 1957, 31, No 2, 463-469.

Abstract: Two series of experiments were carried out: 1/
with a constant content of MnSO_4 (I) and variable
amounts of lithium ferrocyanide (II), and with a
constant content of II and variable amounts of I.
It was found that if the ratio $[\text{Mn}^{2+}] : [\text{Fe}(\text{CN})_6^{4-}]$
(α) in the solution was greater than 2.8, the same
ratio in the precipitate (β) has the approximately

Card 1/2

47

BASINSKI, A.

POLAND / Analytical Chemistry. Analysis of
Inorganic Properties.

E

Abs Jour: Ref Zhur-Khimiya, No 19, 1958, 64189

Author : Basinski Antoni, Kuik Marian

Inst : Not given

Title : Ampere-Metrical Titration of Cadmium by Ferro-
Cyanide of Lithium.

Orig Pub: Roczn. chem., 1957, 31, No 2, 669-675

Abstract: The possibility was studied of ampere-metrical
titration of Cd^{2+} by ferro-cyanide Li (I) and
the reverse. It was established that the presence
of the ions K^+ and NH_4^+ influences the composition
of the sediment that forms, but that against a

Card 1/2

BASINSKI, A

SCIENCE

PERIODICAL: ROCZNIKI CHEMII, Vol. 31, No. 2, 1957

BASINSKI, A. S. Minc and L. Stolarczyk's Elementy fizykochemii koloidow
(Elements of Colloidal Physicochemistry); a book review. p. 752

Monthly List of East European Accessions (EEAI) LC Vol. 8, No. 4
April 1959, Unclass

POLAND / Physical Chemistry. Electrochemistry.

B

Abs Jour: Ref Zhur-Khimiya, No 21, 1958, 70205.

Author : ~~Basinski~~, A. Szymanski, W.

Inst : Not given.

Title : Determination of the Solubility of Cadmium Ferrocyanide in Water by Measuring Specific Electrical Conductance of a Saturated Solution at 25° C

Orig Pub: Roczn. chem., 1958, 32, No 1, 23 - 27.

Abstract: The specific electrical conductance χ of a saturated aqueous solution of $\text{Cd}_2\text{Fe}(\text{CN})_6$ (I) at 25°C has been measured; $\chi = (3.85 \pm 0.08) \cdot 10^{-6}$. Assuming I to be a strong electrolyte the authors calculated the solubility of I according to Onsager's equation to be $5.35 \cdot 10^{-6}$ mole/l.

Card 1/1

BASINSKI, A.

7
Solubility of cobaltous ferrocyanide in water. Antoni
Basinski and Witold Szczepka (Univ. Torun, Poland).
Rozprawy Chem. 33, 283-8 (1960) (English summary).—The
solub. of $\text{Co}_2\text{Fe}(\text{CN})_6$ in water at 20° detd. polarographically
is 3.00×10^{-4} , and detd. colorimetrically 3.5×10^{-4}
moles/l. A. Kreglewski

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BASINSKI, A.

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The solubility of cobalt ferrocyanide in water determined by specific conductivity measurements of a saturated solution at 25°. Antoni Basinski, Wojciech Szymanski, and Teresa Betto (Univ. Torun, Poland). *Roczniki Chem.* 33, 289-91 (1959) (English summary).—Solubility of $\text{Co}_2\text{Fe}(\text{CN})_6$ detd. by sp.-cond. measurements of the satd. soln. at 25°, is 2.55×10^{-5} moles/l. A. Krezlewski

4530

JK

BASINSKI, Antoni; LANGO, Mieczyslaw

Potentiometric titration of thorium ions with potassium
ferrocyanide. Chem anal 4 no.4:691-696 '59. (HEAI 9:6)

1. Zaklad Chemii Fizycznej Uniwersytetu M.Kopernika, Torun.
(Ions) (Potassium ferrocyanide)
(Thorium) (Sodium) (Lithium)

BASINSKI, Antoni; POCZOPKO, Stanislaw

Binodal curve of the ternary system $\text{CdSO}_4 - (\text{CH}_3)_2\text{CO} - \text{H}_2\text{O}$. *Rocz chemii*
33 no.4/5:1109-1115 '59. (EEAI 9:9)

1. Katedra Chemii Fizycznej Uniwersytetu M.Kopernika, Torun.
(Cadmium sulfate) (Acetone) (Water)
(Systems (Chemistry))

BASINSKI, Antoni; CZERNIAWSKI, Marian

Investigations on colloidal properties of vegetable tannins. III.
The fractionation of vegetable tannins. Roczniki chemii 33 no.4/5:
1141-1152 '59. (EKAI 9:9)

1. Zaklad Chemii Fizycznej Uniwersytetu M.Kopernika, Torun.
(Tannins) (Acetone) (Benzene)

BASINSKI, Antoni: CZERNIAWSKI, Marian

Investigations on colloidal properties of vegetable tannins. IV.
Measurements of average molecular weights of some vegetable tannide
extracts. Roczn chemii 33 no.6:1407-1414 '59. (EEAI 9:9)

1. Katedra Chemii Fizycznej Uniwersytetu M.Kopernika, Torun.
(Tannins)

BASINSKI, Antoni; PASGRETA, Maria

On the solubility of manganese ferrocyanide in water and in water solutions. Roczniki chemii 34 no.1:41-45 '60. (EEAI 10:9)

1. Department of Physical Chemistry, Copernicus University, Torun.

(Manganese ferrocyanide) (Water) (Solutions)

BASINSKI, Antoni; ROZWADOWSKI, Michal

On the solubility of nickel ferrocyanide in water and water solutions.
Rocz chemii 34 no.1:47-52 '60. (KEAI 10:9)

1. Department of Physical Chemistry, Copernicus University, Torun.

(Nickel ferrocyanide) (Water) (Solutions)

BASINSKI, Antoni; MUCHA, Janina

The solubility of zinc ferrocyanide in water and in water solutions.
Rocz chemii 34 no.3/4:811-816 '60. (EEAI 10:3)

1. Katedra Chemii Fizycznej Uniwersytetu M.Kopernika, Torun.
(Zinc ferrocyanides)

BASINSKI, Antoni; SKRAGA, Jan

Investigation of properties of tannins derived from lignin. I. The
fractionation of lignin tanning substances. Roczniki chemii 34 no.3/4:
1047-1059 '60. (EEAI 10:3)

1. Katedra Chemii Fizycznej Uniwersytetu M.Kopernika, Torun.
(Tannins) (Lignin)

BASINSKI, Antoni; POZOPKO, Stanislaw

The phenomenon of formation of two liquid phases in some ternary systems in terms of hydration numbers. Roczniki chemii 34 no.3/4:1061-1070 '60. (EAI 10:3)

1. Katedra Chemii Fizycznej Uniwersytetu M.Kopernika, Torun
(Solutions) (Hydration) (Systems (Chemistry))

BASINSKI, Antoni; SKRAGA, Jan

Investigation of properties of tannins derived from lignin. II. The influence of fractionating on the content of tanning substances in tanning lignin preparations. Roczniki chemii 34 no.5:1397-1408 '60.
(KEAI 10:9)

1. Department of Physical Chemistry, M. Copernicus University, Torun.

(Lignin) (Tanning)

BASINSKI, Antoni; ROJEK, Zdzislaw

Investigations on the catalytic properties of cobalt powder. I. The chemisorption of hydrogen on cobalt powder. II. The chemisorption of carbon monoxide on cobalt powder. Roczniki chemii 35 no.4:1083-1101 '61.

1. Department of Physical Chemistry, N. Copernicus University, Torun.

BASINSKI, Antoni; SIEROCKA, Michalina; CYBULSKA, Jadwiga

Studies on the mechanism of the purification of silver halogenide sols by means of ion exchanges. III. The purification of the hydrosol of silver iodide by the static method. Roczniki chemii 35 no.4:1117-1126 '61.

1. Department of Physical Chemistry, N. Copernicus University, Torun.

BASINSKI, Antoni; NAREBSKA, Anna

Determination of acetone and diacetone alcohol in the mixture by ultraviolet absorption method. Roczniki chemii 35 no.4:1131-1141 '61.

1. Department of Physical Chemistry, N. Copernicus University, Torun.

BASINSKI, Antoni; NAREBSKA, Anna

Determination of acetone and diacetone alcohol in presence of mesityl oxide by ultraviolet absorption method. Rocz chemii 35 no.5:1381-1389 '61.

1. Department of Physical Chemistry, M. Copernicus University, Torun.

NAREBSKA, Anna; BASINSKI, Antoni

Catalytic action of anion exchange resins on aldol condensation of acetone and the decomposition of diacetone alcohol.

I. The general mechanism and kinetic equations of the reaction.

II. The condensation kinetics of acetone and the decomposition of diacetone alcohol on amberlit IRA-~~900~~ in OH form.

Rocz chemii 35 no.6:1673-1696 '61.

1. Department of Physical Chemistry, N. Copernicus University, Torun.

BASINSKI, Antoni; LATOWSKI, Tadeusz

On the photochemical properties of halogen derivatives of aniline. I. Photolysis of the iodine derivatives of aniline $J-C_6H_4-R$ ($R=NH_2$, $-NH-CO-CH_3$ and $-NH_2$) in solutions of methanol. Roczniki chemii 36 no.3:513-520 '62.

1. Katedra Chemii Fizycznej, Uniwersytet im. M. Kopernika, Torun i Katedra Chemii Fizycznej, Wyższa Szkoła Pedagogiczna, Gdansk.

S/081/62/000/022/008/088
B177/B186

AUTHORS: Basiński, Antoni, Sierocka, Michalina, Jakubowska, Barbara

TITLE: Investigation into the mechanism of the purification of chlorous silver salts by means of ionites. Part IV. The purification of an AgBr hydrosol obtained by the reaction

$$\text{HBr} + \text{AgNO}_3 = \text{AgBr} + \text{HNO}_3$$

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 22, 1962, 71, abstract 22B502 (Roczn. chem., v. 36, no. 3, 1962, 521-530 [Pol.; summaries in Russ. and Eng.])

TEXT: Dynamic and static methods have been evolved for purifying the negative hydrosol (Hs) of AgBr. It was discovered during the investigation that strong- and medium-basic anionites (e.g. Imac A-17) are unsuitable for purifying Hs of AgBr, as they cause coagulation of the dispersed phase. The weakly basic anionite Lewatit MJH was employed. The optimum flow-rate in the dynamic method was 4 ml/min, the depth of the layer was measured from 3.1 to 11.5 cm. For AgBr in the dispersed

Card 1/2

Investigation into the mechanism ...

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B177/B186

phase at concentrations ≤ 20 mmol/l, the static method of purifying the
Hs can be used. For Part II, see RZhKhim, 1962, 8B553. [Abstracter's
note: Complete translation.]

Card 2/2

3/058/63/000/001/062/120
A160/A101

AUTHORS: Basifski, Antoni, Latowski, Tadeusz

TITLE: The photochemical properties of halogen-derivative anilines. I.
A. photolysis of iodine-derivative anilines $J - C_6H_4 - R$
($R = NH_2$, $-NH - CO - CH_3$ and NH_3^+) in methanol solution ..

PERIODICAL: Referativnyy zhurnal, Fizika, no. 1, 1963, 61, abstract 1D434
("Roczn. chem.", no. 3, 1962, 36, 513 - 520, Polish; summaries
in Russian, English and German)

TEXT: An investigation was carried out of the rate of the photochemical
process of separating iodine from the benzene ring in the presence of the follow-
ing groups: $-NH_2$, $-NH - CO - CH_3$ and $-NH_3$ in methyl alcohol solutions. An
attempt was made to compare the results of the photolysis with the ultraviolet
absorption spectra. ✓

[Abstracter's note: Complete translation]

Card 1/1

BASINSKI, Antoni; KWIATKOWSKI, Edmund

Crystal violet complexes of chlorides and oxychlorides of some elements in solution. I. Optical studies on the antimony pentachloride - crystal violet - chloroform system. *Rocz chemii* 36 no.2:195-202 '62.

1. Department of Physical Chemistry, Copernicus University, Torun, and Department of Physical Chemistry, Normal School, Gdansk.

8/081/63/000/002/015/088
B193/B102

AUTHORS: Basinski, Antoni, Kwiatkowski, Edmund

TITLE: Complexes of crystalline violet with chlorides and oxychlorides of various elements in solution. I. Optical study of the system antimony pentachloride - crystalline violet - chloroform

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 2, 1963, 109, abstract 2V38 (Rozn. chem., v. 36, no. 2, 1962, 195-202 [Pol.; summaries in Russ. and Eng.])

TEXT: Complex formation in chloroform solutions of $SbCl_5$ (I) and crystalline violet (II) was studied spectrophotometrically using methods of continuous variation of molar ratios at $\lambda = 4130$ and $6800 \text{ \AA}$. It was established that complexes form with molar ratios I : II equal to 1, 2 and $1/3$. Complex formation is possible also with molar ratio $1/2$.
[Abstracter's note: Complete translation.]

Card 1/1

S/081/63/000/001/024/061
B144/B186

AUTHORS: Basiński, Antoni, Latowski, Tadeusz

TITLE: Photochemical properties of halogen derivatives of aniline.
I. Photolysis of iodine derivatives of aniline $\text{C}_6\text{H}_4\text{R}$
($\text{R} = \text{NH}_2$, NHCOCH_3 , and NH_3^+) in solutions of methyl alcohol

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 1, 1963, 90, abstract
1B626 (Roczn. chem., v. 36, no. 3, 1962, 513-520 [Pol.;
summaries in Russ., Eng., and Germ.])

TEXT: The photolysis rate of $\text{C}_6\text{H}_4\text{R}$, where $\text{R} = \text{NH}_2$ (I), NHCOCH_3 (II) and NH_3^+ , under the effect of Hg lamps was studied in CH_3OH solutions and in $\text{CH}_3\text{OH}-\text{H}_2\text{SO}_4$ mixtures. For para- and meta-I the photolysis rate is almost the same, for ortho-I it is 24% higher. From the UV spectra it is evident that the photolysis rate increases when the energy of the main electron transition (K band) of the isomers of I and II decreases. This energy reduction is greater, the higher the polarity of the molecule.

Card 1/2

POCZOPKO, Stanislaw; BASINSKI, Antoni; TORSKI, Zbigniew; MANKOWSKA, Wieslawa

Some properties of concentrated solutions of the system
 $\text{MgSO}_4 - \text{Co}(\text{NH}_2)_2 - \text{H}_2\text{O}$. Roczn chemii 36 no.5:947-952 '62.

1. Department of Physical Chemistry, N. Copernicus University,
Torun.

BASINSKI, Antoni; SZYMANSKI, Wojciech; DUDZIK, Malgorzata; KRYGIER, Alicja.

Determination of the solubility of cobalt ferrocyanide and
manganese ferrocyanide by the tracer method using ^{60}Co and
 ^{54}Mn . *Rocz chemii* 36 no.7/8:1139-1142 '62.

1. Department of Physical Chemistry, N.Copernicus University,
Torun.